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$$\text{CH}_4 (\text{g}) + 2\text{H}_2\text{S} (\text{g}) \rightleftharpoons \text{CS}_2 (\text{g}) + 4\text{H}_2 (\text{g}); \Delta H = +232.5 \text{ kJ} \dots$$

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$10^{-3}$  and  $10^3$ . Chemical  
equilibrium (definition) The  
state reached when the  
concentration of reactants

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and products remain constant over time. What is the general, reversible reaction for equilibrium? (With letters)  $aA + bB \rightleftharpoons cC + dD$ . Equilibrium equation (general with letters)  $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ .

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state from where neither the

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26 D5 Chemical equilibrium  
is said to be dynamic  
because A. the reaction

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proceeds quickly. B. the mass of the reactants is decreasing. C. the macroscopic properties are constant. D. both forward and reverse reactions are occurring. 27 D5 Equilibrium is a dynamic process because

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the A. macroscopic  
properties are not changing.

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\*\*See your instructor to  
find out what questions you  
should complete for review.

2.3 (a) State in which  
direction the equilibrium  
will shift for each of the  
following, when there are

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The given reaction equation

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is:  $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$   
 $\Delta H = -191 \text{ kJ}$   
 $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g}) \quad \Delta H = -191 \text{ kJ}$ . The equilibrium constant is a ratio of the forward and reverse ...

~~Consider the following~~

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1) Give a balanced chemical equation which describes this equilibrium. 2) Explain what influence the removal of  $H_2$  from the container will have on the

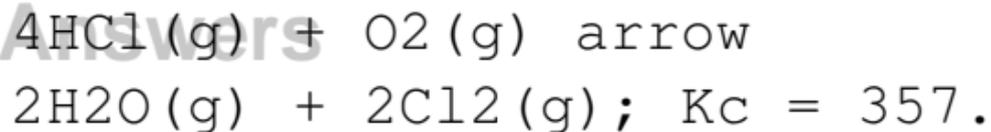
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equilibrium. 3) Explain what effect...

~~You are given a container in which  $H_2$  (g) ...  
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Solved: Consider the following equilibrium:

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An equilibrium mixture of the system was found to...

~~Consider the following equilibrium:  $4\text{HCl}(\text{g})$  —~~  
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Answers Changing the temperature of a reaction at equilibrium alters both the equilibrium. (12) and the equilibrium position. When a reaction is. (13), which means it releases energy, lowering the temperature shifts the

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equilibrium to the  
(14) because the forward  
reaction liberates heat and  
removes the (15).

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