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An acid is a compound that donates a proton.  
After the proton leaves, the remaining  
compound (with one less proton) is the  
conjugate base. Example:  $\text{NH}_4^+ + \text{H}_2\text{O} \rightarrow \text{NH}_3 + \text{H}_3\text{O}^+$   
Acid =  $\text{NH}_4^+$  Conjugate base =  $\text{NH}_3$

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reactants is the Brønsted-Lowry acid, and which is the Brønsted-Lowry base? Explain your answer.  $\text{H}_2\text{CO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{HCO}_3^-(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$ . The  $\text{H}_2\text{CO}_3$  is the Brønsted-Lowry acid because it donates a proton to the  $\text{H}_2\text{O}$ . The  $\text{H}_2\text{O}$  is the Brønsted-Lowry base because it accepts the proton. b.

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acid 21. hypobromous acid 22.  $H_2S$  23.  $HNO_3$   
24.  $H_3PO_3$  25.  $HClO_4$  26.  $NH_3$  27. acidic:  
 $HCl$  and  $NH_4^+$  basic:  $NH_3$  and  $Cl^-$  28.  $H_2O$  and  
 $OH^-$   $NH_3$  and  $NH_4^+$  29. proton donors:  $H_2O$  and  
 $NH_4^+$  proton acceptors:  $OH^-$  and  $NH_3$  30.  
 $HCl(aq)$   $NaOH(aq)$  ?  $NaCl(aq)$   $H_2O(l)$  31.  $H_3O^+$   
 $Cl^-(aq)$   $Na^+(aq)$

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DR. DENNIS ENGLIN earned his bachelor's from Westmont College, his master of science from California State University, and his EdD from the University of Southern California. He enjoys teaching animal biology, vertebrate biology, wildlife biology, organismic biology, and astronomy at The Master's University. His professional memberships include the Creation Research Society, the American Fisheries Association, Southern California Academy of Sciences, Yellowstone Association, and Au Sable Institute of Environmental Studies.

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